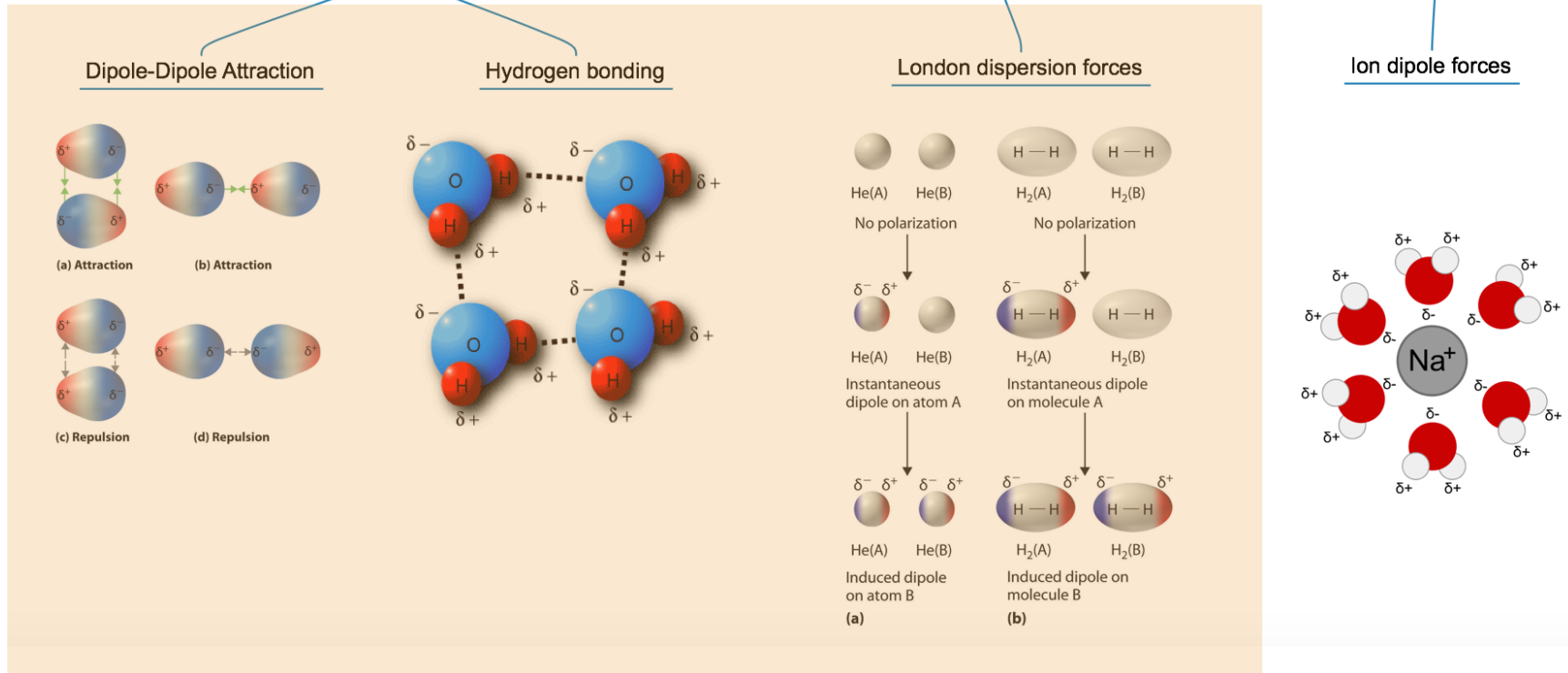
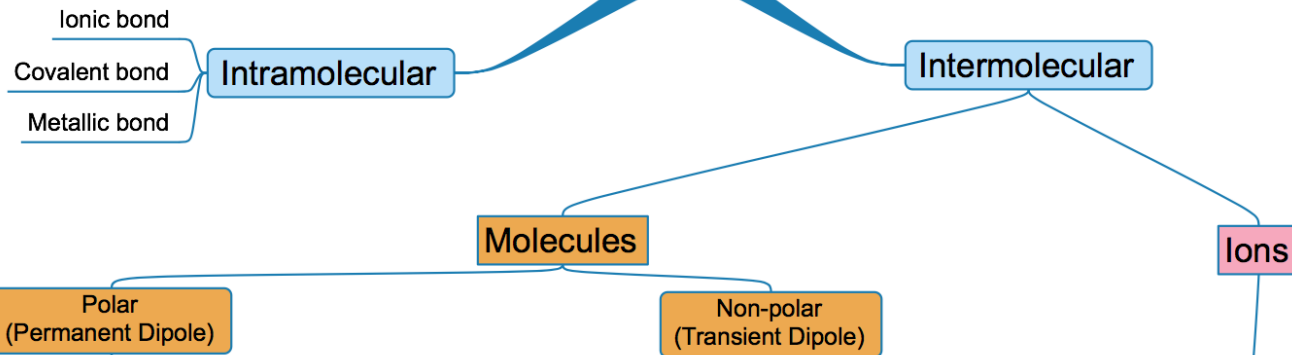


Bonding forces



Van der Waals interactions

Bond Type	Dissociation Energy (kJ mol ⁻¹)
Ionic (halides)	400
Covalent (halides)	150 - 400
<i>Van der Waals forces</i>	
Hydrogen bonds	10 - 16
Dipole-dipole	0.5 - 2.0
London dispersion	< 1.0

Van der Waals forces increase with:

- increasing **polarity** (polar molecules attract and repel more strongly than non-polar molecules)
- increasing **polarizability** (ability of molecule to displace electrons and become polar - more e- generally lead to more polarizability)